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[Solution Stoichiometry Worksheet](#) Solve the following solutions Stoichiometry problems: 1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate? $2 \text{ AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{ KNO}_3(\text{aq})$ 0.150 L AgNO_3 0.500 moles AgNO_3 1 moles Ag_2CrO_4 331.74 g $\text{Ag}_2\text{CrO}_4 = 12.4 \text{ g Ag}_2\text{CrO}_4$ 1 L 2 moles ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left Stoichiometry worksheet answers chapter 9. Show all your work in the space provided. 1. 4 Stoichiometry worksheet answers chapter 9. 5 mol The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally ...

Chapter 10 The Mole Worksheet Answers

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Solution Stoichiometry . Name _____ CHEMISTRY 110 . last first ... Answer _____ 3] 60.5 mL of HNO_3 are required to react with 25.0 mL of a 1.00 M Barium hydroxide solution: $2 \cdot \text{HNO}_3(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow 2 \cdot \text{H}_2\text{O}(\text{l}) + \text{Ba}(\text{NO}_3)_2(\text{aq})$ (BALANCED) Find the . Molarity. of the nitric acid solution . 25.0ml $\text{Ba}(\text{OH})_2$ X . 1L 10. 3. ml X. 1L 1.00mol $\text{Ba}(\text{OH})_2$. X 2mol HNO_3 1mol $\text{Ba}(\text{OH})_2 = 0.0500\text{mol HNO}_3$. M ...

WORKSHEET 13 Name - Cerritos College

CHM 130 Stoichiometry Worksheet The following flow chart may help you work stoichiometry problems. Remember to pay careful attention to what you are given, and what you are trying to find. 1.

Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide: $\text{C}_6\text{H}_{12}\text{O}_6(\text{s}) \rightarrow 2\text{C}_2\text{H}_5\text{OH}(\text{l}) + 2\text{CO}_2(\text{g})$ A. Calculate the mass of ethanol produced if ...

CHM 130 Stoichiometry Worksheet

Stoichiometry Practice Worksheet The correct answer is A. $\text{H}_2\text{C}_2\text{O}_2$ is the limiting reagent: $(4.5\text{mol H}_2\text{C}_2\text{O}_4) / (2\text{mol MnSO}_4 / \text{mol H}_2\text{C}_2\text{O}_4)$ According to the equation $2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$, sodium reacts with water. Stoichiometry Flashcards | Quizlet Worksheet 14 1 Worksheet #14 Limiting Reagents 1.

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