

## Chapter 11 Review Gases Section 2 Answers Modern Chemistry

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~~FSc Physics Book1, CH 11, LEC 1: Pressure of Gases Chapter 10 (Gases) - Part 1 review questions chapter 11 Class 10(Physics) World Climate \u0026amp; Climate Change - Chapter 12 Geography NCERT Class 11 Class 10th-Physics-Chapter 11-Sound-Exercise-Review Questions Chapter 11 Review Gases Section~~

~~SECTION 1 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. b Pressure — orce For a constant force, when the surface area is tripled the surface area pressure is (a) doubled. as much. (c) ripld. 7-0 (d) unchanged. Rank the following pressures in increasing order. (c) 76 torr (a) 50 kPa O, OO\ictbv-x~~

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If a gas and a liquid are the same temperature and pressure, diffusion occurs much faster in the gas because. A. there are more elastic collisions between the particles in a gas. B. gases are more compressible. C. the particles move faster in a gas and there is a greater distance between them.

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Section 11.4 Dalton's Law of Partial Pressures Goals To describe the properties of mixtures of gases. To describe calculations that deal with mixtures of gases. In the real world, gases are usually mixtures. This section describes how mixing gases affects the properties of the resulting mixture.

*Chapter 11 - Gases*

462 Chapter 11 Gases Discovering the Relationships Between Properties If we want to explain why a weather balloon carrying instruments into the upper atmosphere expands as it rises, we need to consider changes in the properties of the gases (pressure, volume, temperature, or number of gas particles) inside and outside the balloon.

*Chapter 11 Gases - An Introduction to Chemistry*

CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. d, c, a, b Rank the following pressures in increasing order. (a) 50 kPa (c) 76 torr (b) 2 atm (d) 100 N/m<sup>2</sup> 3.

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this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 Key Te r m s pressure newton barometer millimeters of mercury

*SECTION 11.1 Gases and Pressure - Pickford High School*

CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same

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