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\u0026 Moles

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Chemistry- Stoichiometry 5:

Summary Flowchart Honors

Chemistry, 5/4/2020,

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Stoichiometry

Stoichiometry Worksheet #1

Answers. Stoichiometry

Worksheet #1 Answers 1.

Given the following

equation: $2 \text{ C } 4\text{H } 10 + 13 \text{ O}$

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2 \rightarrow 8 CO_2 + 10 H_2O , show what the following molar ratios should be. a. C 4H 10 / O 2 b. O 2 / CO_2 c. O 2 / H_2O d. C 4H 10 / CO_2 e. C 4H 10 / H_2O 2. Given the following equation: $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ a.

Honors Chemistry

Stoichiometry Practice 1

Answers

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Stoichiometry Honors

Chemistry Extra

Stoichiometry Problems 1.

Silver nitrate reacts with

barium chloride to form

silver chloride and barium

nitrate. a. Write and

balance the chemical

equation. $2 \text{AgNO}_3 + \text{BaCl}_2$

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2. $\text{AgCl} + \text{Ba}(\text{NO}_3)_2$ b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many

Honors Chemistry
Stoichiometry Problems 1
Answers ...

Dr. Rodriguez-Reyes
Chemistry Honors

Stoichiometry problems I.
Answer the questions for the following reaction: $\text{Na (s)} + \text{Cl}_2 \rightarrow \text{NaCl (s)}$

1. How many moles of Na are needed to react with 12.50 mole Cl_2 ?
2. How many moles of Na are needed to produce 56.79 mole of NaCl?
3. How many moles of Cl_2 are needed to react with 33.50 mole Na?
- 4.

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Stoichiometry Problems 1

2018 Stoichiometry
worksheet.docx - Dr
Rodriguez-Reyes ...

Homework–Solving

Stoichiometry Problems Name

___ ANSWERS _____ If the
statement is true, write

“true”. If it is false,
change the underlined word
or words to make it true.

Write your answer on the
line provided. ___ TRUE _____

1. The major types of
stoichiometry problems are
mass-mass, mass-volume, and
volume-volume.

Homework–Solving

Stoichiometry Problems

Honors Chemistry Practice

Worksheet – Stoichiometry.

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1. How many moles of oxygen are consumed when 96.7 moles of hydrogen sulfide gas are burned, producing sulfur dioxide and water vapor in the process? 2. If 3.70×10^{23} molecules of oxygen react with excess benzene (C_6H_6), how many grams of water can be produced? 3.

Honors Chemistry Practice
Worksheet – Stoichiometry
Honors Chemistry Extra
Stoichiometry Problems 1.
Silver nitrate reacts with
barium chloride to form
silver chloride and barium
nitrate. a. Write and
balance the chemical
equation. $2 AgNO_3 + BaCl_2 \rightarrow$
 $2 AgCl + Ba(NO_3)_2$ b. If

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39.02 grams of barium
chloride are reacted in an
excess of silver nitrate,
how many

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HW4 Solutions-Molarity-
Stoichiometry WS 1-14
Answers Page 1 Page 2 HW5
Activity 5-8: 1-8 Answers
Page 1 Page 2 ... Answers to
Chemistry Final Review .
Honors Chemistry
Assignments. Acids and Bases
TEST Wed June 7 HW1 (5/30)
Definitions - handed in ...

Bader, Mr. K. - Science /
Honors Chem Homework
The Stoichiometry of Alka-
Seltzer. This lab will count

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as your Honors Project. You will submit a typed, formal lab report, including all pre-lab and post-lab assignments. It will count as a formal grade worth 100 points. Alka-Seltzer is one of the world's best-known antacids. Its main function is to absorb excess stomach acid (HCl).

The Stoichiometry of Alka Seltzer

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work and ask questions when
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Chapter 12 SG 12.1

Introduction to
Stoichiometry ...

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Mr. Christopherson /
Stoichiometry
Chemistry I-Honors.
Stoichiometry P.S.#2. A
student performs a double
replacement reaction by
mixing 500.0 ml of a 0.228 M
solution of lithium
carbonate with 370.0 ml of a
0.352 M solution of
iron(III) chloride. The
student collects the
precipitate, and finds that
9.98 grams of precipitate
are recovered. 1. Write the
net ionic equation.

Chemistry I-Honors
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for students who have

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demonstrated strong ability
in previous science courses
Unit 8 stoichiometry test
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reactions, stoichiometry,
gases, solutions, reaction
kinetics, equilibrium, acid-
base theory, oxidation-
reduction, and ...

Unit 8 Stoichiometry Test
Review Answer Key
Honors Stoichiometry
Problems 1) $10+ \text{CO} \rightarrow \text{CO}_2$
Co. a). 2.00g of carbon
monoxide reacted with
dioxide, calculate

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the theoretical yield of 1
b). If 3.179 of t, was
experimentally (actually)
produced, calculate the
percent yield of 1, 2). CH_0 ,
 NH , + H_2 CH_2N + H_0 a).

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Problems 1) $10+$ CO CO \gg Co

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(stoichiometry) of 1 mol NO
: 3 mol NO₂ (or it takes 3
mol NO₂ to make 1 mol
NO)... we can set up the
following relationship
(using dimensional
analysis): 1.2 moles NO₂ x
1 mole NO / 3 moles NO₂ =
0.4 moles NO formed. 82).
Write a correctly balanced
equation for the reaction
taking place: 2NO(g) + O₂
(g) ==> 2NO₂ (g)

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